

Study Guide The Mole

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Study Guide The Mole

The mole is the most common unit used to express the quantity of a chemical substance. For all solids, liquids, and gases, you can convert mass to moles (or moles to mass). For gases, you should memorize the following conversion of volume to moles (or moles to volume): at 0°C and 1 atmosphere pressure (known as standard temperature and pressure or STP), one mole of any gas occupies approximately 22.4 liters.

The Mole Unit - CliffsNotes Study Guides

Chapter 10 Study Guide: The Mole Matching Match each item with the correct statement below. a. molar volume b. molar mass c. atomic mass ____ 1. the number of grams of an element that is numerically equal to the atomic mass of the element in amu ____ 2. the mass of a mole of any element or compound ____ 3. the volume occupied by a mole of any ...

Chapter 10 Study Guide: The Mole

5.65 x 1024 atoms divided by 6.02 x 1023 = 9.39 mole x 78. 96 mass = 741 grams. How many atoms of gallium are in 2.850 g of gallium. 2,850 g divided by 69.72 = 40.98 moles x 6.02 x 1023 = 2.46 x 1025. Determine the mass in grams of 0.0458 mol sulfur. 0.0458 mol x 32.066 g = 1.47 grams.

Study Guide the Mole Flashcards | Quizlet

1 mole compound=6.022 X 10²³ molecules 1 mole element=Av# atoms† 1 mole ionic compound=Av# formula units Atom to amu; grams to moles=Periodic table atoms/molecules to moles=Av#

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Study Guide - Chapter 10 - The Mole Section 10.1 Measuring Matter 1. pair 2. 5 3. dozen 4. gross 5. 200 6. ream 7. 6,000,000,000 8. 0.5 mol 9. 6.02 1023 10. four moles 11. 6.02 10 Cu atoms23 1 mol Cu 12. 4 23 4 1 mol CH 6.02 10 molecules CH 13. 23 1 mol Xe 6.02 10 molecules Xe 14. 23 2 2 6.02 10 molecules F 1 mol F Section 10.2 Mass and the ...

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The Mole Summary - eNotes.com

Mole is a physical count of particles (e.g. atoms, molecules, ions, and compounds). The quantity is just a count and is not linked to a universal mass-mole conversion factor. To illustrate, a mole...

What is the meaning of 'moles' in chemistry? | Study.com

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moles =9.47×1020moles × mole 6.022×1023 atoms = 1.57×10−3 moles m o l e s = 9.47 × 10 20 m o l e s × m o l e 6.022 × 10 23 a t o m s = 1.57 × 10 − 3 m o l e s .

Calculate the number of moles of each atom ... - study.com

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Chapter 10 Study Guide The Mole Answer Key

moles Molar Mass = Cu + N² + O⁹ + 6^H = 241.6 g/mol . 48.4 g CuN3W. x . mol CuN3W = 0.20 mol Cu(NO3).3H2O 241.6 g CuN3W Mole-Mole Stoichiometry. Given the reaction 4 Fe + 3 O 2 → 2 Fe 2O 3. 8) Given 64 moles of Fe and excess O 2, how many moles of Fe 2O 3 can be made? 64 mol Fe x 2 mol Fe2O3 = 32 mol Fe2O3 4 mol Fe . 9) Given 48 moles of O2 and excess Fe, how many moles of Fe 2O 3 can be made?

Name Period Stoichiometry - The Mole Study Guide

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